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Part 3 Chapter 16 Review Acid Base

Section 16.1 – ACIDS AND BASES: A BRIEF REVIEW •

Acids and bases were first recognized by the properties of their aqueous solutions. o For example, acids turn litmus red,

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whereas bases turn litmus blue.

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BASES. Table 16.1. 1: General
Properties of Acids and Bases.
produce a piercing pain in a wound.

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give a slippery feel. taste sour.
taste bitter. are colorless when
placed in phenolphthalein (an
indicator). are pink when placed in
phenolphthalein (an indicator). are
red on blue litmus paper (a pH
indicator).

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Weak Bases; 16.8 Relationship
Between K_a and K_b ; 16.9 Acid-
Base Properties of Salt Solutions;
16.10 Acid-Base Behavior and
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Chapter 16: Acid and Base Review
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~PLEASE DO NOT WRITE ON
THIS WORKSHEET~ 1. What two
substances are always produced

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by a neutralization reaction? a. acid and a base b. water and a base c. water and an acid d. water and a salt 2.

Chapter 16: Leader: Acid and Base
Review

Chapter 16 Acids and Bases 1.

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Acids were recognized primarily from their sour taste. Bases were recognized from their bitter taste and slippery feel on skin. 2. In the Arrhenius definition, an acid is a substance that produces hydrogen ions (H^+) when dissolved in water, whereas a base is a

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substance that produces hydroxide ions (OH^-) in

Chapter 16 Acids and Bases

9/15/12 1 1 Chapter 16 Acids and

Bases 16.1 Acids and Bases: A

Brief Review 16.2 Brønsted-

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Autoionization of Water 16.4 The
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Acid-Base Behavior and Chemical
Structure 16.11 Lewis Acids and

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Bases Ch. 16 Mastering

Chemistry; Due September 26,
2012 2 Overview The Arrhenius
definition is the narrowest view of
...

Chapter 16 - Acid Base
Equilibrium - Chapter 16 Acid-

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16.1 Acids and Bases: A Brief Review

- Acids taste sour and cause certain dyes to change color.
- Bases taste bitter and feel soapy.
- Arrhenius concept of acids and bases:
- An acid is a substance that, when dissolved in

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water, increases the concentration of H^+ ions. • Example: HCl is an acid. • An Arrhenius base is a substance that, when dissolved in water, increases the concentration of OH^- ions.

AP Chemistry— CHAPTER 16

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STUDY GUIDE Acid-Base Equilibrium

Acids and Bases Acid and Base
Strength In any acid-base reaction,
the equilibrium will favor the
reaction that moves the proton to
the stronger base. $\text{HCl (aq)} + \text{H}_2\text{O(l)}$
 $\rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{Cl}^-(\text{aq})$

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H₂O is a much stronger base than Cl⁻, so the equilibrium lies so far to the right K is not measured ($K \gg 1$).

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16.10: Acid-Base Behavior and
Chemical Structure Inductive
effects and charge delocalization

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significantly influence the acidity or basicity of a compound. The acid – base strength of a molecule depends strongly on its structure. The weaker the A – H or B – H⁺ bond, the more likely it is to dissociate to form an (H^+) ion.

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CHAPTER 16 – Acid-Base

Equilibria Section 16.1 – Acids and

Bases: A Brief Review (a) Define

an acid and a base, according to

the Arrhenius definition. acid =

base = (b) Write the products of

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each chemical reaction below,
which involves the dissociation of
each reactant into aqueous ions.

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This video explains the concepts from your packet on Chapter 16 (Acid-Base Equilibria), which can be found here:

<https://goo.gl/MV7sAR> Section
16.1: Acids an...

Chapter 16 Acid-Base Equilibria -

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Chapter 16 – Acid Base Equilibria.

16.1 Acids & Bases: A Brief

Review. Arrhenius acids and

bases: acid: an H^+ donor $HA \rightleftharpoons H^+$

$A^-(aq)$ base: an OH^-

donor $MOH \rightleftharpoons M^+ + OH^-(aq)$

Brønsted Lowry acids and bases:

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acid: an H^+ donor $HA \rightleftharpoons H^+(aq) + A^-(aq)$
base: an H^+ acceptor
HB

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tutorial focuses on acids and bases and buffer solutions. It shows you how to calculate the pH and pOH of the solution. It cont...

Ka Kb Kw pH pOH pKa pKb H+
OH- Calculations - Acids ...

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2: Acid-Base Titration and pH
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Learners examine an acid-base
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